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Chapter 8 - Basic
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Valency | Covalent bond

|Chemical bonding (part

3)| 9th science chapter 8

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Covalent Bonding

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Vocabulary 8

unshared pair. a pair of valence electrons that is not shared between atoms (lone pair) double covalent bond. bond involving shared pairs of electrons. triple covalent bond. a bond formed by sharing three pairs of electrons. coordinate covalent bond. a covalent bond in which one atom

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contributes both
bonding electrons.

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Chemistry Chapter 8
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Bonding ...*

When H forms a bond
with H O to form the
hydronium ion H O ,
this bond is called a
coordinate covalent
bond because a. both
bonding electrons come
from the oxygen atom.

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b. it forms an especially strong bond. c. the electrons are equally shared. d. the oxygen no longer has eight valence electrons.

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a pair of valence
electrons not shared
between atoms. double
covalent bond. a bond

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Vocabulary

that involves two shared pairs of electrons. triple covalent bond. a bond formed by sharing three pairs of electrons.

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coordinate covalent bond. a covalent bond in which one atom contributes both bonding electrons. polyatomic ion.

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Bonding Flashcards ...

-A covalent Bond is the sharing of valence

electrons-A molecule is what forms when 2 or more atoms bond

covalently o Vs ionic

compound-Covalent

bonds generally occur

between elements that

are near each other on

the periodic table o In

particular, between

nonmetal elements

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Electronegativity- the attraction that an atom has to a shared pair of electrons-Elements in the upper right corner of the periodic table have a very strong attraction for shared pairs of electrons-This means that they ...

*Chem chapter 8 lecture
\u2013 covalent
bonding.docx - Chem ...
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Chapter 8 Covalent
Bonding Crossword
Down : 2) A covalent
bond in which the
electrons are shared
equally between the two
atoms. 3) A bond
formed when two atoms
share a pair of electrons
(2 electrons).

*Vocabulary Crossword
Puzzle: Chapter 8
Covalent Bonding ...
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single covalent bond:
atoms joined together by
sharing a pair of (2)
electrons: double
covalent bond: atoms
joined together by
sharing two pairs of (4)
electrons: triple covalent
bond: atoms joined
together by sharing
three pairs of (6)
electrons: molecular
compound: compound
composed of a neutral

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Chapter 8
group of atoms joined
by covalent bonds:
molecular formula

Bonding
Quia - Chapter 8

"Covalent Bonding"

Review Answer
Chapter 8 Covalent

Key
Bonding. chemical

bond. covalent bond.

molecule. electron dot

diagram. ???the force

that holds two atoms

together. ???the

chemical bond that

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Vocabulary
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results from sharing
valence electro.... ??a
molecule is formed
when two or more
atoms bond covalently.
???used to show
valence electrons of
atoms.

*chapter 8 covalent
bonding Flashcards and
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When sharing of
electrons occurs, the

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attachment between atoms that results is called a. covalent bond.

When such an attachment is formed, bond dissociation energy is released, and the process is.

exothermic. When two or more atoms bond by means of electron sharing, the resulting particle is called a. molecule.

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COVALENT

BONDING Class 8.2

8.2 8.2 8.2 8.3 8.3 8.3

8.3 195 Chapter Quiz

Choose the best answer
and write its letter on
the line. 1. A bond in
which each atom
contributes two
electrons is a. a double

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covalent bond. b. an ionic bond. c. a polar covalent bond. d. a coordinate covalent bond. 2. The electron dot structure for hydrogen sulfide, H₂S, is b.

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8 Review Covalent
Bonding Vocabulary 1.
Covalent bond - a bond
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formed by sharing of electrons between atoms

2. Molecule - a neutral group of atoms joined together by covalent bonds 3.

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Single Covalent Bonds
When only one pair of electrons is shared, such as in a hydrogen

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molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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Types of covalent

bonds. A. Single

covalent bond– 1 shared

pair of valence

electrons: 2 dots, or a

single dash, represent 2

electrons that are

simultaneously being

attracted by, or “shared”

by, the nuclei of two

neighboring atoms. 1.

The formula in the

center is a type of

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structural formula called
a “Lewis dot structural
formula”.

Chapter 8: Covalent Bonding

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Covalent Bonding

Single Covalent Bonds

When only one pair of
electrons is shared, such
as in a hydrogen
molecule, it is a single
covalent bond. The

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relative
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