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Ion. Consists of a single atom with a
positive or negative charge resulting from
the loss of gain of one or more valence
electrons and behaves as a unit, yet it still
has a charge. (+ or -). Determined by the

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PLAY. Monatomic ion. - single atom with
a positive or negative charge resulting
from the loss or gain of electrons,
respectively.

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Section 9.4 – Naming and Writing

Formulas for Acids and Bases. An acid is
a compound that produces H^+ ions when it
dissolves in water. The formula for an acid
normally starts with and H. When naming
acids, you should first determine whether
or not the acid contains oxygen.

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83 SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268–270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming

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Binary Molecular Compounds (pages 268–269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

~~Name Date Class CHEMICAL NAMES
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SECTION 9.2 NAMING AND WRITING
FORMULAS FOR IONIC COMPOUNDS

I. Write the formulas for these binary ionic
compounds. c. potassium iodide 2. write
the formulas for the compounds formed c.

b. f. 3. Name the following binary ionic
compounds. sodium sulfide g. CuCl : h.

snc. l. a. MnO_2 c. d. SrBr_2 221 Chapter 9
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9.1 Naming Ions 9.2 Naming and Writing

Formulas for Ionic Compounds 9.3

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Naming and Writing Formulas for
Molecular Compounds 9.4 Naming and
Writing Formulas for Acids and Bases 9.5
The Laws Governing How Compounds
Form

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Chapter 9 “Chemical Names and
Formulas” ... in parenthesis after the name
of the metal (Table 9.2, p.255) Predicting
Ionic Charges Some of the post-transition
elements also have more than one possible
oxidation state. Tin (II) = Sn^{2+} Lead (II) =
 Pb^{2+} Tin (IV) = Sn^{4+} Lead (IV) = Pb^{4+}

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CHAPTER 9 Assessment 42. a. 2– b. 1+ c.

1– d. 3+ 43. a. 2+ b. 2+ c. 3+ d. 1+

44. a. barium ion b. iodide ion c. silver ion d.

mercury(II) ion 45. cyanide, CN^- and

hydroxide, OH^- 46. a. hydroxide ion b.

lead(IV) ion c. sulfate ion d. oxide ion 47.

zero; A compound is electrically neutral.

48. The symbols for the cation and

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Naming Ions. 1. What is the charge on the

ion typically formed by each element? a.

oxygen c. sodium e. nickel, two electrons

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lost . b. iodine d. aluminum f. magnesium.

2. How many electrons does the neutral atom gain or lose when each ion forms? a.

Cr³⁺ c. Li⁺ e. Cl⁻ b. P₃ ...

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activities; ... In samples of any chemical compound, the masses of the elements are always in the same proportions any atom or group of atoms that has a negative charge ... place cation name first followed by the anion name: naming polyatomic ...

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