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This text was written with an aim to provide the beginner with a reliable and understandable guide for study in the teacher's absence. Except where it would needlessly overburden the student, the subject is presented in a mathematically rigorous way. In spite of this, no mathematics beyond the elementary calculus is required.

A leading book for 80 years. Silbey's Physical Chemistry features exceptionally clear explanations of the concepts and methods of physical chemistry for students who have had a year of calculus and a year of physics. The basic theory of chemistry is presented from the viewpoint of academic physical chemists, but the many practical applications of physical chemistry are integrated throughout the text. The problems in the text also reflect a skillful blend of theory and practical applications. This text is ideally suited for a standard undergraduate physical chemistry course taken by chemistry, chemical engineering, and biochemistry majors in their junior or senior year.

Mathematics for Physical Chemistry, Third Edition, is the ideal text for students and physical chemists who want to sharpen their mathematics skills. It can help prepare the reader for an undergraduate course, serve as a supplementary text for use during a course, or serve as a reference for graduate students and practicing chemists. The text concentrates on applications instead of theory, and, although the emphasis is on physical chemistry, it can also be useful in general chemistry courses. The Third Edition includes new exercises in each chapter that provide practice in a technique immediately after discussion or example and encourage self-study. The first ten chapters are constructed around a sequence of mathematical topics, with a gradual progression into more advanced material. The final chapter discusses mathematical topics needed in the analysis of experimental data. Numerous examples and problems interspersed throughout the presentations Each extensive chapter contains a preview, objectives, and summary Includes topics not found in similar books, such as a review of general algebra and an introduction to group theory Provides chemistry specific instruction without the distraction of abstract concepts or theoretical issues in pure mathematics

The first IUPAC Manual of Symbols and Terminology for Physicochemical Quantities and Units (the Green Book) of which this is the direct successor, was published in 1969, with the object of 'securing clarity and precision, and wider agreement in the use of symbols, by chemists in different countries, among physicists, chemists and engineers, and by editors of scientific journals'. Subsequent revisions have taken account of many developments in the field, culminating in the major extension and revision represented by the 1988 edition under the simplified title Quantities, Units and Symbols in Physical Chemistry. This 2007, Third Edition, is a further revision of the material which reflects the experience of the contributors with the previous editions. The book has been systematically brought up to date and new sections have been added. It strives to improve the exchange of scientific information among the readers in different disciplines and across different nations. In a rapidly expanding volume of scientific literature where each discipline has a tendency to retreat into its own jargon this book attempts to provide a readable compilation of widely used terms and symbols from many sources together with brief understandable definitions. This is the definitive guide for scientists and organizations working across a multitude of disciplines requiring internationally approved nomenclature.

Written by Ira Levine, the Student Solutions Manual contains the worked-out solutions to all of the problems in the text. The purpose of the manual is help the student learn physical chemistry and as an incentive to work problems, not as a way to avoid working problems.

This book covers various metallurgical topics, viz. roasting of sulfide minerals, matte smelting, slag, reduction of oxides and reduction smelting, interfacial phenomena, steelmaking, secondary steelmaking, role of halides in extraction of metals, refining, hydrometallurgy and electrometallurgy. Each chapter is illustrated with appropriate examples of applications of the technique in extraction of some common, reactive, rare or refractory metal together with worked out problems explaining the principle of the operation.

This book is a physical chemistry textbook that presents theessentials of physical chemistry as a logical sequence from itsmost modest beginning to contemporary research topics. Many bookscurrently on the market focus on the problem sets with a cursorytreatment of the conceptual background and theoretical material,whereas this book is concerned only with the conceptual developmentof the subject. Comprised of 19 chapters, the book willaddress ideal gas laws, real gases, the thermodynamics of simplesystems, thermochemistry, entropy and the second law, the Gibbsfree energy, equilibrium, statistical approaches to thermodynamics,the phase rule, chemical kinetics, liquids and solids, solutionchemistry, conductivity, electrochemical cells, atomic theory, wavemechanics of simple systems, molecular orbital theory, experimentaldetermination of molecular structure, and photochemistry and thetheory of chemical kinetics.

Thermodynamics Problem Solving in Physical Chemistry: Study Guide and Map is an innovative and unique workbook that guides physical chemistry students through the decision-making process to assess a problem situation, create appropriate solutions, and gain confidence through practice solving physical chemistry problems. The workbook includes six major sections with 20 - 30 solved problems in each section that span from easy, single objective questions to difficult, multistep analysis problems. Each section of the workbook contains key points that highlight major features of the topic to remind students of what they need to apply to solve problems in the topic area. Key Features: Includes a visual map that shows how all the equations used in thermodynamics are connected and how they are derived from the three major energy laws. Acts as a guide in deriving the correct solution to a problem. Illustrates the questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry Can be used as a stand-alone product for review of Thermodynamics questions for major tests.

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